

CHEM 36
General Chemistry
Quiz #6

March 29, 2002

Name: _____

1. To 10.0 mL of a 0.10 M Acetic Acid solution, 5.0 mL of a 0.10 M NaOH solution is added. Classify the resulting solution by circling one of the following (but remember, you must show your work to get any credit!):

Weak Acid (HAc)
Buffer
Weak Base (Ac⁻)
Strong Base (excess OH⁻)

2. At the equivalence point of a titration of acetic acid with NaOH, the pH is:

7.00 < 7.00 > 7.00 (Circle your answer)

Briefly explain how you arrived at your answer.

3. The pH of a solution prepared by dissolving solid NH₄Cl in water will be:

7.00 < 7.00 > 7.00 (Circle your answer)

Briefly explain how you arrived at your answer.